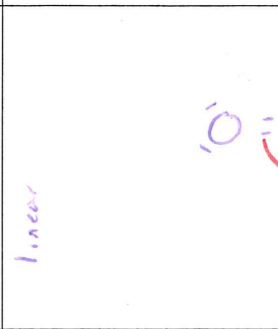
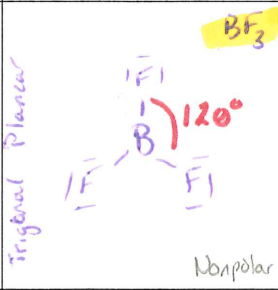

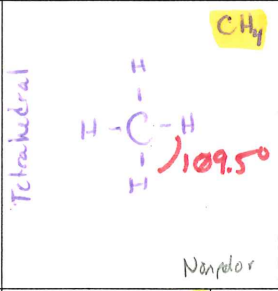
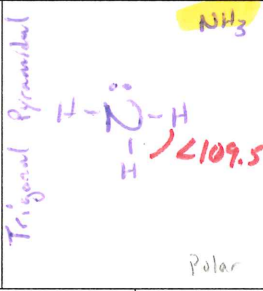
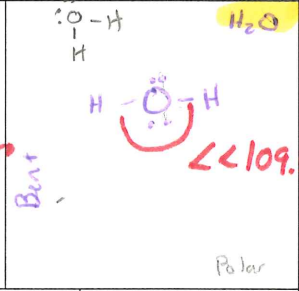
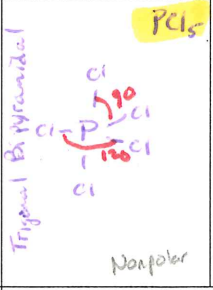
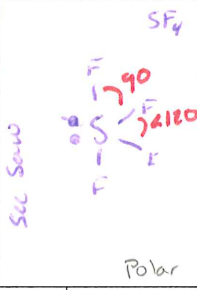
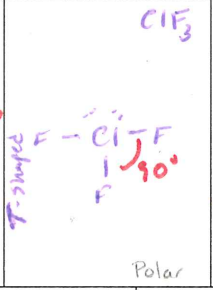
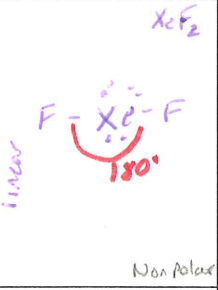
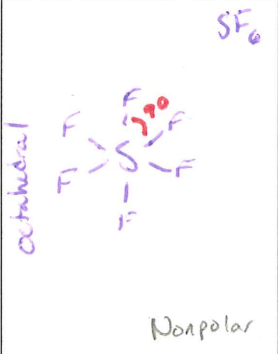
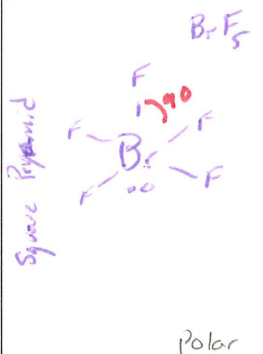
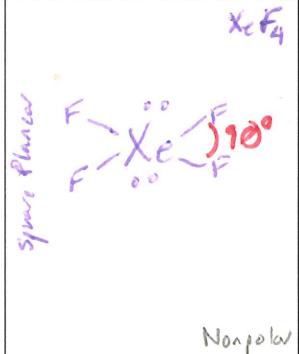


HYBRIDIZATION	REGIONS OF ELECTRON DENSITY	GEOMETRIES ANGLES POLARITIES			
sp 180°	2	CO_2 		Remember that in molecules where the outside atoms are different, shapes that tend to be nonpolar usually become polar	
sp^2 120°	3	BF_3 	SO_2 	Remember to count the number of regions of electron density. Lone pairs, double and triple bonds are all only one region	
sp^3 109.5°	4	CH_4 	NH_3 	H_2O 	
dsp^3 90° 120°	5	PCl_5 	SF_4 	ClF_3 	XeF_2 
d^2sp^3 90°	6	SF_6 	B_2F_5 	XeF_4 	

Lone pairs and atoms with higher electronegativities cause a bond to have a dipole moment towards them
Lone pairs have more energy and thus occupy more space, they repel bonded pairs therefore decreasing bond angles.